UKA TARSADIA UNIVERSITY
Maliba Pharmacy College
B. Pharm 3rd Semester Internal Examination 2013 (*Mid-Sem II*)
030020304- Pharmaceutical Analysis 1

Time: 10:30 a.m. to 12:30 p.m.   Max. Marks: **40**
Date: 26/11/2013

Instructions:
- ➢ Attempt any **FIVE** questions.
- ➢ Each question carries **08** marks.
- ➢ Make suitable assumption whenever necessary.
- ➢ Figures to the right indicate full marks.

Q.1	**A)**	Discuss applications of complexometric titrations	04
	B)	50.0 ml of a solution which is 0.01 M in calcium cation and buffered at pH 10.0 is titrated with 0.01 M EDTA solution. Calculate values of pCa after addition of 0 ml, 10.0 ml, 50.0 ml and 60.0 ml of titrant. (The K_{eff} value for complex of calcium = 1.8×10^{10})	04
Q.2	**A)**	Define non-aqueous titration and discuss differentiating and leveling effect of solvent	04
	B)	Classify and define different types of solvent with suitable examples.	04
Q.3	**A)**	Enlist different factors affecting solubility of slightly soluble salts and discuss any two in detail.	04
	B)	50.0 ml of 0.1 M sodium chloride solution is titrated with 0.1 M silver nitrate solution. Calculate values of pCl after addition of 0 ml, 10.0 ml, 50.0 ml and 60.0 ml of titrant.(K_{sp} for silver chloride = 1×10^{-10})	04
Q.4	**A)**	Write a short note on indicators used in precipitation titration	04
	B)	Discuss Von Weimarn's theory of relative super saturation for gravimetric method	04
Q.5	**A)**	Discuss different types of precipitations in gravimetric method of analysis.	04
	B)	Write a short note on organic precipitant used for gravimetric analysis	04
Q.6	**A)**	Enlist different type of Redox titration. Discuss any one in detail.	04
	B)	Differentiate Iodometry and Iodimetry	04
Q.7	**A)**	Describe preparation and standardization of 0.1M sodium thiosulphate solution	04
	B)	Enlist different end point detection method used in redox titration. Discuss them.	04